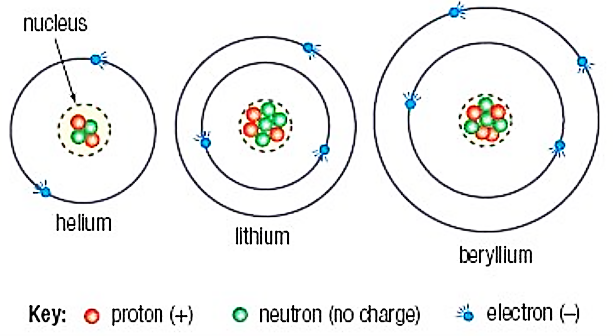
**Atoms – Protons, Neutrons, and Electrons**

Anythiong you can touch and feel are made from atoms everything. There are no exceptions, even the air is made up of atoms – in air the atoms are spaced apart a lot so air is easy to move through – but they are still there. When you feel the wind you are feeling atoms move from one place to another.

 An atom is the very smallest particle which exists of an element. All of the atoms of any one element (say oxygen) are identical. Oxygen gas is made from trillions and trillions of identical oxygen atoms. There are only 90 naturally occurring different elements. They are listed in the periodic table along with some human made atoms. Everything in the universe is made from combinations of these 90 naturally occurring atoms.

However, some of these 90 atoms are very rare, and almost all of the matter you can see is actually made up of only twenty or thirty common atoms. Luckily, atoms can join together in millions of different combinations to make **all** the substances on Earth and beyond.

**Structure of the atom**

Every atom is made of a nucleus consisting of protons and neutrons. The nucleus is surrounded by electrons.

It is the number of protons an atom has which gives it its identity, so for example, all oxygen atoms have exactly 8 protons.

Protons are positively charged and electrons are negatively charged. Neutrons have no charge. This means the nucleus (protons and neutrons) of an atom is positively charged. The negatively charged electrons move around the nucleus and are held in place by their attraction to the positively charged nucleus.

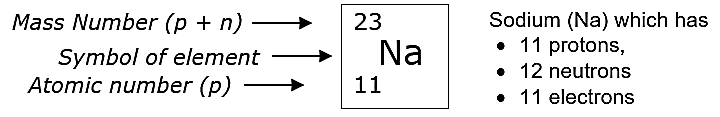
An atom has a neutral overall charge because it usually has the same number of electrons as protons (same number of negative and positive charges).

Protons and neutrons have the same **mass**. Electrons have such a small mass that this can usually be taken as zero.

|  | **Proton** | **Neutron** | **Electron** |
| --- | --- | --- | --- |
| **Charge** | +1 | 0 | -1 |
| **“Relative”**  **Mass** | 1 | 1 | 0.0005 (almost zero) |

Comparing the charge and mass of electrons, protons and neutrons

The **atomic number** is the number of protons in an atom. In a neutral atom (where the protons and electrons are equal in number) the number of electrons is also equal to the atomic number.

The **mass number** (also called the nucleon number) is the total number of protons and neutrons in an atom.

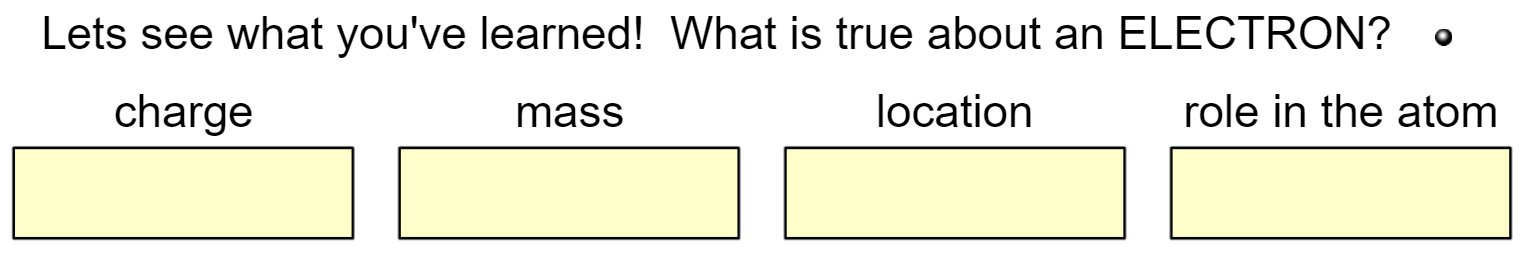
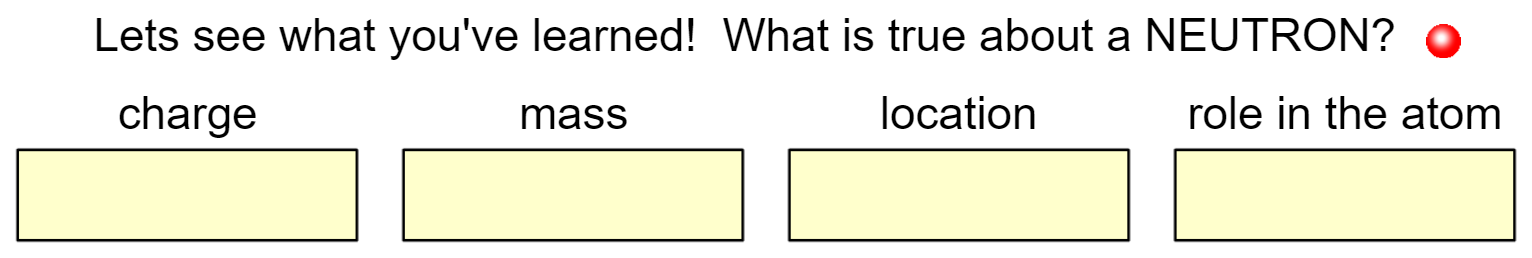
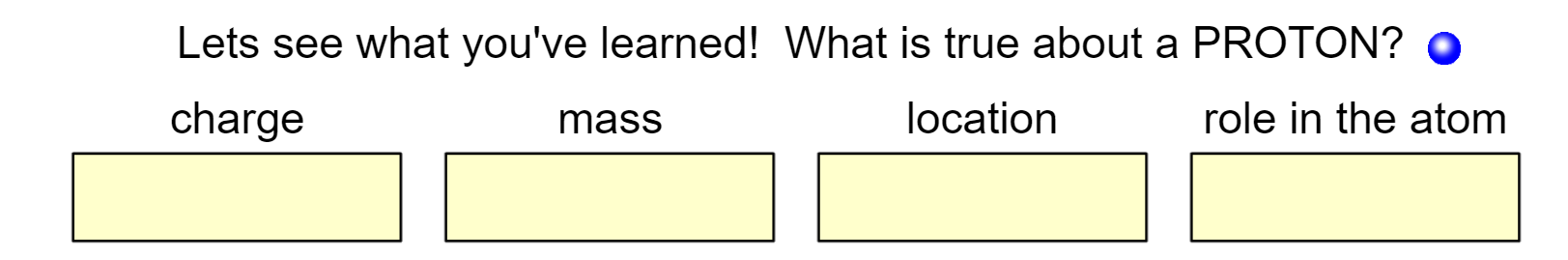
The elements are arranged in the periodic table in lowest to highest (ascending) order of atomic number - so it's easy to find the name or symbol for an atom if you know the atomic number.

**VERY IMPORTANT:** Elements are described mainly by their atomic numbers as this tells you where the atom is in the periodic table - and the number of protons the atom has. A neutral atom (most) has the same number of electrons AND the number of electrons is what determines the chemistry of the element – how it reacts, what it reacts with, and how many bonds it forms. Your teacher can explain this in more detail when you look at a periodic table.

Go to

<https://contrib.pbslearningmedia.org/WGBH/arct15/SimBucket/Simulations/chemthink-atomicstructure/content/index.html>

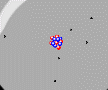
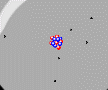
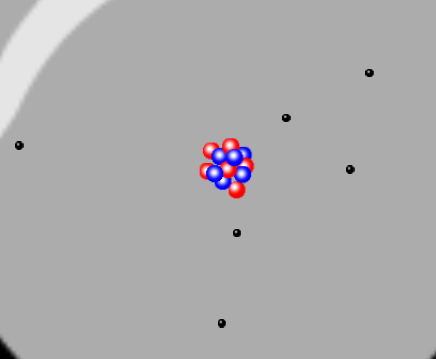
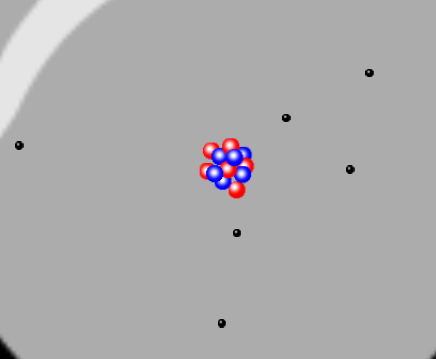
and complete the slide show which has activities for you to complete in the last few slides. Copy your answers into the boxes below



Identify the name of the atom in the picture on the left: \_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Identify the number of protons: \_\_\_\_

Neutrons: \_\_\_\_



Electrons: \_\_\_\_